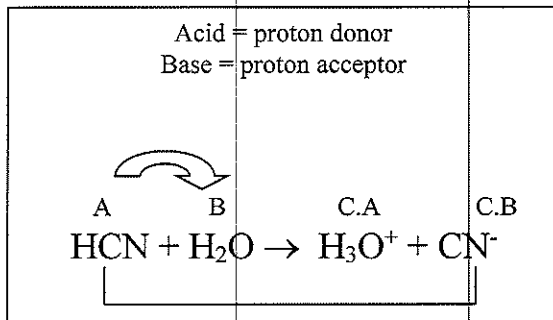


How to reaction produce H_3O^+ and OH^- ions?

- What are the ways in which a reaction can produce H^+ or OH^-
 - Student will be able to characterize the different ways reactions can make H^+ and OH^- are produced.
 - Students will be able to write hydrolysis reactions for acids and bases.



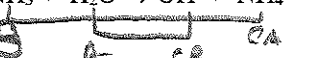
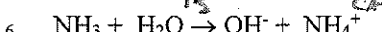
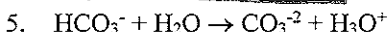
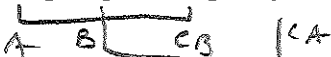
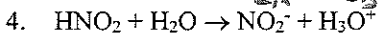
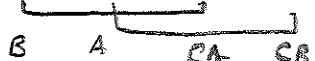
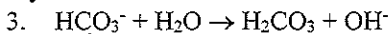
1. How many ways can an acid produce Hydronium Ions? Explain.



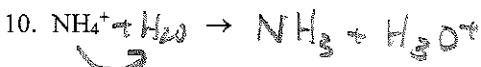
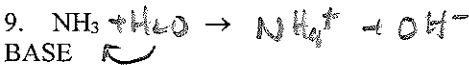
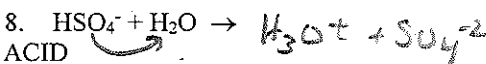
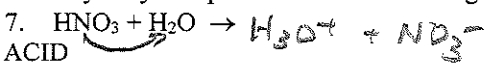
2. How many different ways can base produce hydroxide ions? Explain.



Identify the Brønsted acid-base and the conjugates in each example

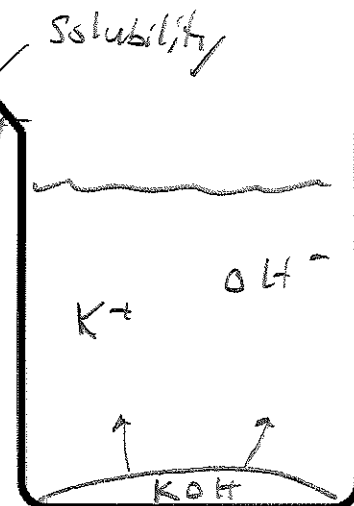


Write the Hydrolysis equations for the following acids and bases

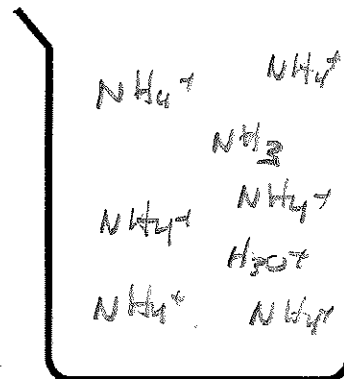


11. Show Beaker A filled with water after having some solid KOH dissolved

12. Show Beaker B filled with NH_4^+ dissolving in water and a small fraction undergoing hydrolysis. (add an anion where appropriate.)



Beaker A



Beaker B