

(#8-1)

What makes something acidic or basic?

- o Student will be able to identify acids and bases and whether a solution is acidic or basic.
- o Student will be able to give names and formulas to acids.

1. Name each of the following binary acids:

- o HCl *Hydrochloric Acid*
- o H₂S *hydrosulfuric Acid*

2. Name each of the following oxyacids:

- o HNO₃ *Nitric Acid*
- o H₂SO₃ *Sulfurous Acid*
- o HClO₃ *Chloric Acid*
- o HNO₂ *Nitrous Acid*

3. Write the formulas for each of the following binary acids:

- o hydrofluoric acid *HF(aq)*
- o hydriodic acid *HI(aq)*

4. Write formulas for each of the following oxyacids:

- o bromic acid *HBrO₃(aq)*
- o chlorous acid *ClO₂(aq)*
- o phosphoric acid *H₃PO₄(aq)*
- o hypochlorous acid *HClO(aq)*

Strong Bases

Write out the formula/name of these strong bases.

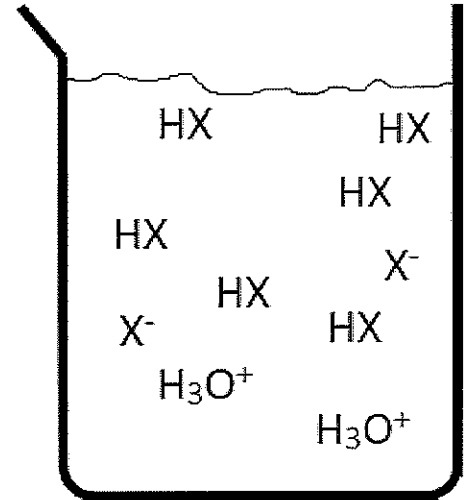
- 5. LiOH *Lithium Hydroxide*
- 6. Calcium Hydroxide *Ca(OH)₂*
- 7. KOH *Potassium hydroxide*
- 8. Sodium Hydroxide *NaOH*
- 9. Mg(OH)₂ *magnesium hydroxide*

Write out the formula/names of these weak bases.

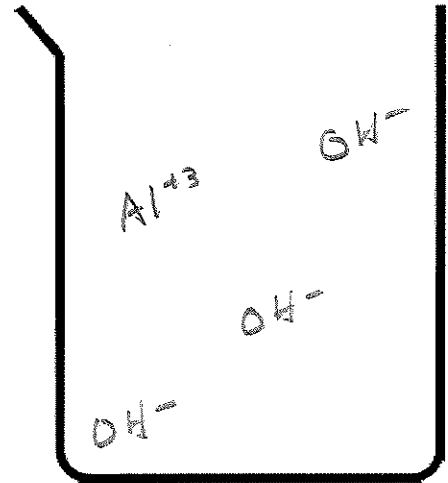
- 10. Iron (II) Hydroxide *Fe(OH)₂*
- 11. Ammonia *NH₃*
- 12. Al(OH)₃ *Aluminium hydroxide*
- 13. Be(OH)₂ *Beryllium hydroxide*
- 14. Cu(OH)₂ *Copper II hydroxide*

15. Beaker A, is an Acid/Base/ neutral? Justify?
Contains H⁺ or H₃O⁺

16. In Beaker B, Draw a solution of #12.



Beaker A



Beaker B