

(#8-1)

What makes something acidic or basic?

- o Student will be able to identify acids and bases and whether a solution is acidic or basic.
- o Student will be able to give names and formulas to acids.

1. Name each of the following binary acids:

- o HCl *Hydrochloric acid*
- o H₂S *Hydro sulfuric acid*

2. Name each of the following oxyacids:

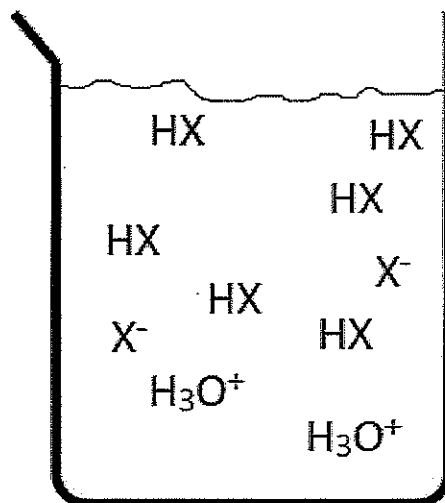
- o HNO₃ *Nitric Acid*
- o H₂SO₃ *Sulfurous acid*
- o HClO₃ *chloric acid*
- o HNO₂ *Nitrous acid*

3. Write the formulas for each of the following binary acids:

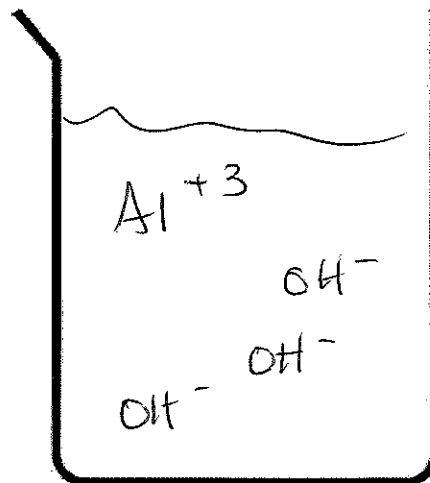
- o hydrofluoric acid *HF aq*
- o hydriodic acid *HI aq*

4. Write formulas for each of the following oxyacids:

- o bromic acid *HBrO₃ aq*
- o chlorous acid *HClO₂ aq*
- o phosphoric acid *H₃PO₄ aq*
- o hypochlorous acid *HClO aq*



Beaker A



Beaker B

Strong Bases

Write out the formula/name of these strong bases.

- 5. LiOH *Lithium Hydroxide*
- 6. Calcium Hydroxide *Calcium Hydroxide*
- 7. KOH *Potassium Hydroxide*
- 8. Sodium Hydroxide *Sodium Hydroxide*
- 9. Mg(OH)₂ *Magnesium Hydroxide*

Write out the formula/names of these weak bases.

- 10. Iron (II) Hydroxide *Fe(OH)₂*
- 11. Ammonia *NH₃*
- 12. Al(OH)₃ *Aluminum Hydroxide*
- 13. Be(OH)₂ *Beryllium Hydroxide*
- 14. Cu(OH)₂ *Copper(II) Hydroxide*

- 15. Beaker A, is an Acid/Base/neutral? Justify?
→ H₃O⁺ is acidic
- 16. In Beaker B, Draw a solution of #12.

