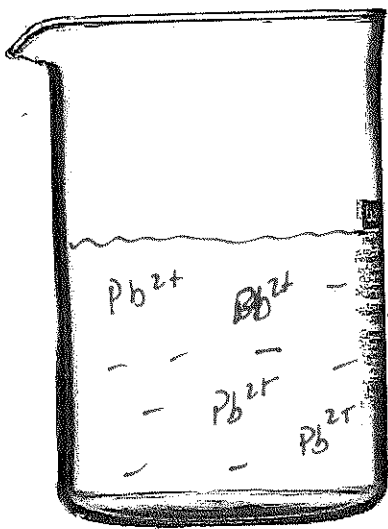
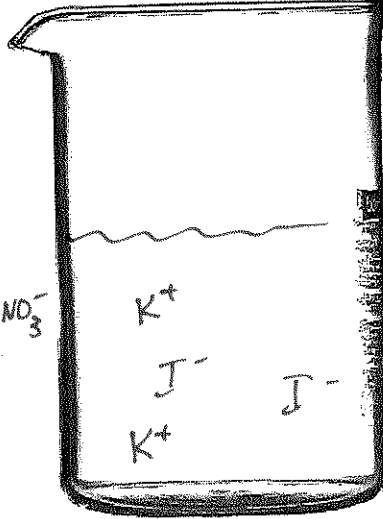


AS

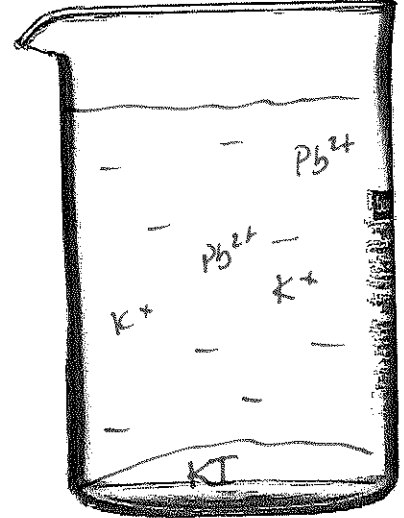
Lead (II) Iodide Demonstration



Before: Pb(NO₃)₂



Before: KI

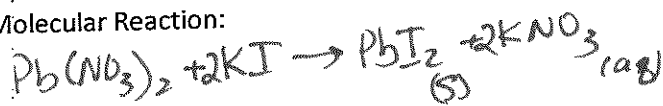


after

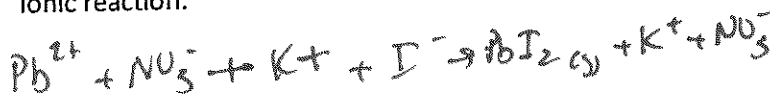
Sketch the beakers before starting.

Excess lead (II) Nitrate is added to KI

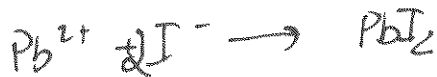
Molecular Reaction:



Ionic reaction:



Net ionic reaction:



Notes on how to model Solubility reactions



Different combinations of different ions have different solubilities

What ions are routinely soluble?

- Na⁺
- K⁺
- NH₄⁺
- NO₃⁻

What are the spectator ions here?

- Do not change!

What does it mean to be in

Excess: still present after rxn stops

Limiting: - Ran out stopping rxn