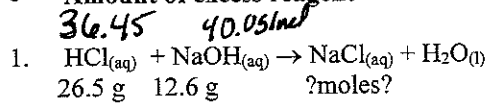


NAME
CHEMISTRY
LIMITING REACTANT & EXCESS

For each of the following reactions determine

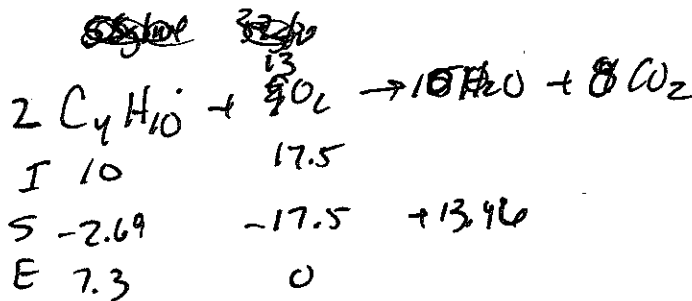
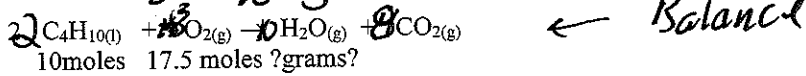
- Balance reaction
- Answer ? in correct units
- The limiting reactant
- Amount of excess reagent



I 0.72 0.315
 S -0.315 -0.315 (+0.315) +0.315
 E 0.405 0

$26.5 \text{ g} \cdot \frac{1 \text{ mol}}{36.45 \text{ g}} = 0.72 \text{ mol}$

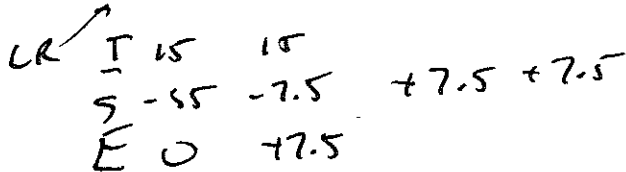
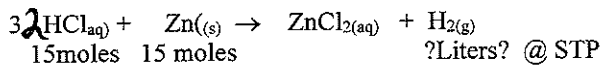
$12.6 \text{ g} \cdot \frac{1 \text{ mol}}{40 \text{ g}} = 0.315 \text{ mol}$



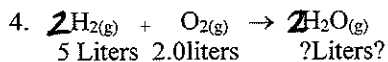
$17.5 \cdot \frac{2}{13} = 2.69$

$17.5 \cdot \frac{14}{13} = 18.96$

$13.96 \cdot \frac{18 \text{ g}}{1 \text{ mol}} = 242 \text{ g}$



$7.5 \cdot \frac{22.4 \text{ L}}{1 \text{ mol}} = 168 \text{ L}$



Use Liters just like moles

