

Section Review

Energy and Rates of Chemical Reactions

USING KEY TERMS

The statements below are false. For each statement, replace the underlined term to make a true statement.

1. An exothermic reaction absorbs energy.

2. The rate of a reaction can be decreased by adding a catalyst.

UNDERSTANDING KEY IDEAS

- _____ 3. Which of the following will not increase the rate of a reaction?

- a. adding a catalyst
- b. increasing the temperature of the reaction
- c. decreasing the concentration of reactants
- d. grinding a solid into powder

4. How does the concentration of a solution affect the rate of reaction?

CRITICAL THINKING

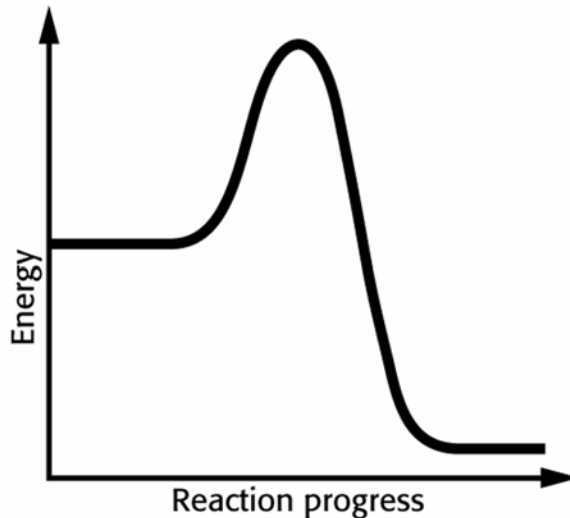
5. **Making Comparisons** Compare exothermic and endothermic reactions.

6. **Applying Concepts** Explain how chewing your food thoroughly can help your body digest food?

Section Review *continued*

INTERPRETING GRAPHICS

Use the diagram below to answer the questions that follow.



7. Does this energy diagram show an exothermic or an endothermic reaction?
How can you tell?

8. A catalyst lowers the amount of activation energy needed to get a reaction started. What do you think the diagram would look like if a catalyst were added?
