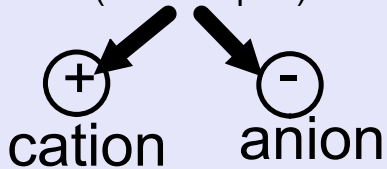


Types of Bonds -Dissolving in water

IONIC

(transfer e⁻)

-dissociates
(breaks apart)



-charges interact
with H bonds of H₂O

COVALENT

(share e⁻)

-no dissociation

Polar

contains
N, O, or F

non-symmetrical



-charged areas
interact with H bonds
of H₂O

NonPolar

same atoms

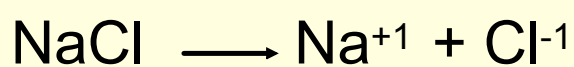


symmetrical

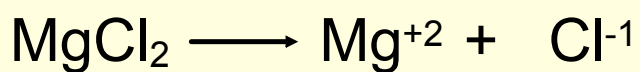
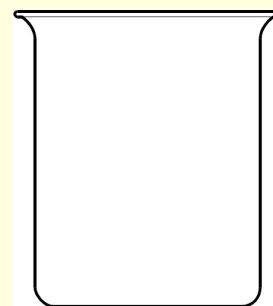



-does not interact with H
bonds of H₂O

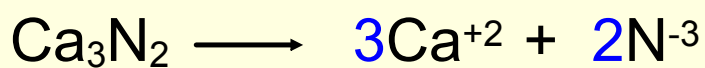
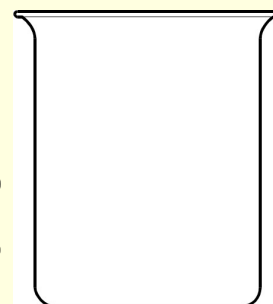
Ionic salts dissolving
dissociation and modeling:





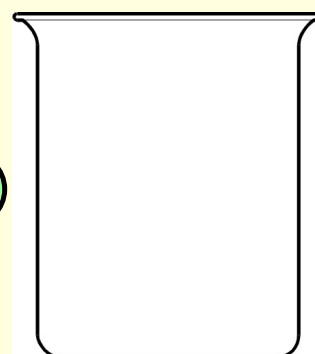
Na⁺¹
Cl⁻¹



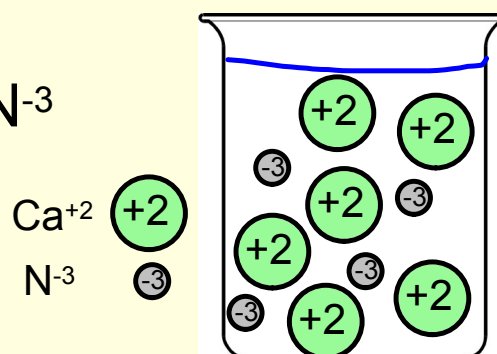
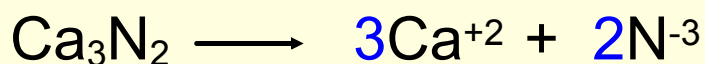
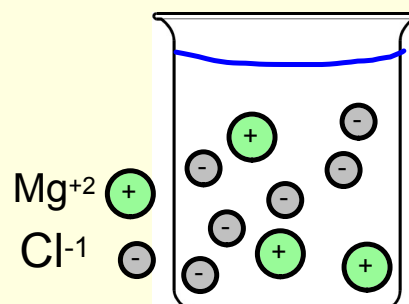
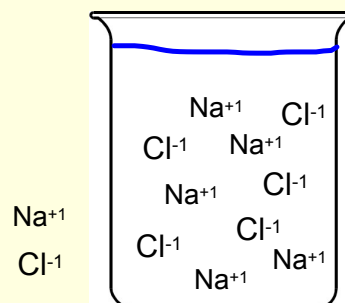
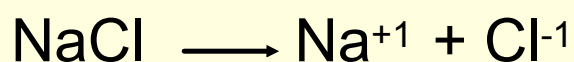
Mg⁺² 
Cl⁻¹ 



Ca⁺² 
N⁻³ 

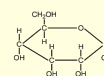
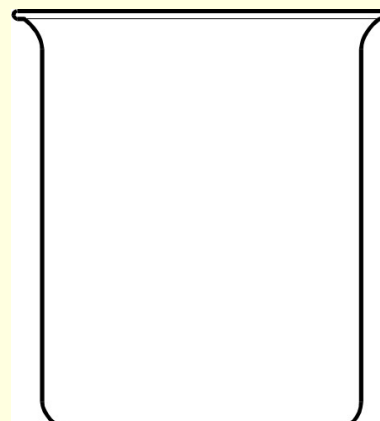
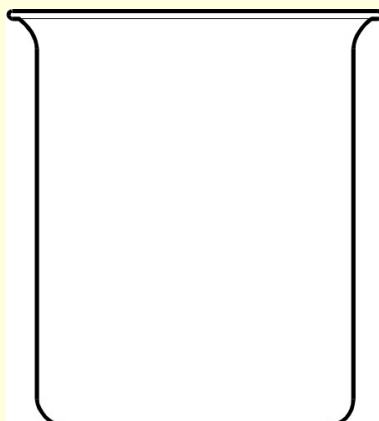
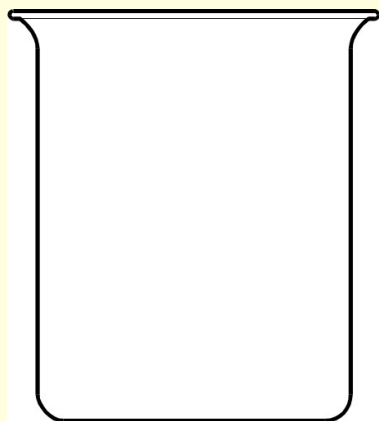
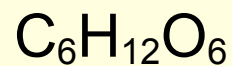
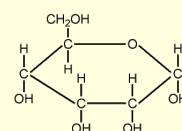


Ionic salts dissolving
dissociation and modeling:



Covalent substances dissolving

no dissociation, molecules do not break apart



Covalent substances dissolving

no dissociation, molecules do not break apart

