

## Copper Conservation Rubric

### Part A-G (10 points)

- Equations or Chemical reactions for each process. Ex: Part G:  $\text{Cu}(\text{SO}_4)_2 + \text{Zn} \rightarrow \text{Cu}(\text{s}) + \text{ZnSO}_4(\text{aq})$
- Type of reaction: Oxidation Reduction
- Picture and Description

### Types of Reactions you will see

- Oxidation Reduction:
  - $\text{AB}_{(\text{aq})} + \text{C}_{(\text{s})} \rightarrow \text{CB}_{(\text{aq})} + \text{A}_{(\text{s})}$
  - There are more complicated versions of this reaction. A typical google search will help to clear up any questions. Or just ask and I will help... only issue is you can't ask me 5 seconds before it is due. Be prepared and get ahead.
- Acid base Neutralization Acid + base  $\rightarrow$  water
  - $\text{H}_3\text{O}^+ + \text{OH}^- \rightarrow 2\text{H}_2\text{O}$
  - This goes for any time you neutralized with acid or base
- Solubility reaction
  - $\text{AB}_{(\text{aq})} + \text{CD}_{(\text{AQ})} \rightarrow \text{AD}_{(\text{aq})} + \text{CB}_{(\text{s})}$
- Decomposition: Single item breaking down. Usually by heat being added.
  - $\text{A} \rightarrow \text{B} + \text{C}$

### Lab packet (5 points)

Pre-lab questions and fully completed lab data packet. So your packet will be getting turned in on Monday. You must have your own packet. Not your partners.

### Post lab quiz (5 points)

This will be completed and you may use your lab packet. Both will be turned in. Not able to retake this.

Any Questions see me!

RESPECT THE CHEMISTRY!